

Molar Mass Review

Name _____ Period _____ Date _____

1) What is the formula and molar mass for Calcium Chloride?

2) What is the formula and molar mass for Iron (II) Bromide?

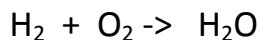
3) What is the formula and molar mass for sulfur trioxide?

4) What is the formula and molar mass for Sodium Sulfate?

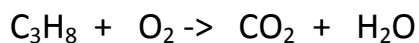
5) What is the formula and molar mass for Nitric Acid?

6) What is the formula and molar mass for Hydrochloric Acid?

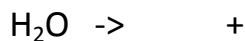
7) Balance this formula:



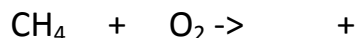
8) Balance this formula



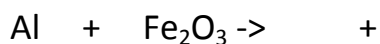
9) Identify the following reaction, predict the products, and balance:



- 10) Identify the following reaction, predict the products, and balance:



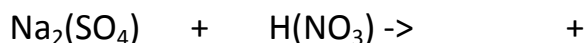
- 11) Identify the following reaction, predict the products, and balance:



- 12) Identify the following reaction, predict the products, and balance:



- 13) Identify the following reaction, predict the products, and balance:



- 14) What are the 3 common states of matter? What symbols are used to identify them?

- 15) Calculate the number of moles of sodium sulfate is needed to produce 10.0 moles of sodium nitrate.

- 16) Calculate the number of moles of oxygen needed to produce 2.50 moles of water.

After Snow Review (KEY)

1) What is the formula for Calcium Chloride? CaCl_2 111.1 g/mol

2) What is the formula for Iron (II) Bromide? FeBr_2 215.7 g/mol

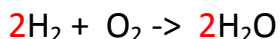
3) What is the formula for Sulfur Trioxide? SO_3 80.1 g/mol

4) What is the formula for Sodium Sulfate? $\text{Na}_2(\text{SO}_4)$ 142.1 g/mol

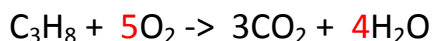
5) What is the formula for Nitric Acid? $\text{H}(\text{NO}_3)$ 63.0 g/mol

6) What is the formula for Hydrochloric Acid? HCl 36.5 g/mol

7) Balance this formula



8) Balance this formula



9) Identify the following reaction and predict the products



10) Identify the following reaction and predict the products



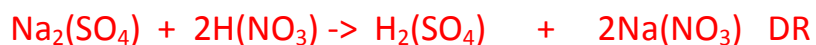
11) Identify the following reaction and predict the products



12) Identify the following reaction and predict the products



13) Identify the following reaction and predict the products



14) What are the 3 common states of matter? What symbols are used to identify them?

Solid, Liquid, Gas (s), (l), and (g) (aq) dissolved in water

Example:

Ice - $\text{H}_2\text{O}_{(s)}$

Water – $\text{H}_2\text{O}_{(l)}$

Steam – $\text{H}_2\text{O}_{(g)}$

15) Calculate the number of moles of sodium sulfate is needed to produce 10.0 moles of sodium nitrate.

5.00 moles of sodium sulfate

16) Calculate the number of moles of oxygen needed to produce 2.50 moles of water.

1.25 moles of O_2